

Pb₃Te₂O₆Br₂

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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{Te}-\text{O}) = 0.015$ Å; R factor = 0.048; wR factor = 0.107; data-to-parameter ratio = 18.8.

Single crystals of the title compound, trilead(II) bis-[tellurate(IV)] dibromide, have been grown under hydrothermal conditions. The structure is isotopic with that of the chloride analogue, $\text{Pb}_3\text{Te}_2\text{O}_6\text{Cl}_2$, and consists of three Pb, two Te, two Br and four O atoms in the asymmetric unit. Except for two of the O atoms, all atoms are located on mirror planes. The $\text{Pb}_3\text{Te}_2\text{O}_6\text{Br}_2$ structure can be described as being built up from $\infty^2[\text{Pb}_3\text{Te}_2\text{O}_6]^{2+}$ layers extending parallel to $(20\bar{1})$ and Br^- anions between the layers. Cohesion of the structure is accomplished through Pb—Br contacts of two of the three lead atoms, leading to highly asymmetric coordination polyhedra. The lone-pair electrons of both Te^{IV} and Pb^{II} atoms are stereochemically active and point towards the anionic halide layers.

Related literature

For reports and structures of other compounds in the system $\text{Pb}X_2\text{—PbO—TeO}_2$, where $X = \text{Br}, \text{Cl}$, see: $\text{Pb}_3\text{TeO}_4X_2$ (Charkin *et al.*, 2006; Porter & Halasyamani, 2003); $\text{Pb}_3\text{Te}_2\text{O}_6X_2$ (Porter & Halasyamani, 2003). The crystal chemistry of oxotellurate(IV) compounds has been reviewed by Dolgikh (1991) and Zemann (1971). For other oxotellurates(IV) prepared under hydrothermal conditions, see: Weil & Stöger (2007, 2008a,b).

Experimental*Crystal data*

$\text{Pb}_3\text{Te}_2\text{O}_6\text{Br}_2$	$V = 1031.3$ (1) Å ³
$M_r = 1132.59$	$Z = 4$
Monoclinic, $C2/m$	Mo $K\alpha$ radiation
$a = 16.9151$ (9) Å	$\mu = 62.14$ mm ⁻¹
$b = 5.6813$ (3) Å	$T = 296$ K
$c = 11.0623$ (6) Å	$0.18 \times 0.10 \times 0.04$ mm
$\beta = 104.046$ (1)°	

Data collection

Bruker APEXII CCD diffractometer	Absorption correction: numerical (<i>HABITUS</i> ; Herrendorf, 1997)
	$T_{\text{min}} = 0.06, T_{\text{max}} = 0.41$

3758 measured reflections
1371 independent reflections

1338 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.030$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.048$
 $wR(F^2) = 0.107$
 $S = 1.15$
1371 reflections

73 parameters
 $\Delta\rho_{\text{max}} = 7.41$ e Å⁻³
 $\Delta\rho_{\text{min}} = -6.41$ e Å⁻³

Table 1

Comparative geometrical parameters (Å) for selected bond lengths in $\text{Pb}_3\text{Te}_2\text{O}_6X_2$ compounds ($X = \text{Br}, \text{Cl}$).

Distance	$X = \text{Br}$ (this work)	$X = \text{Cl}$ (Porter & Halasyamani, 2003)
Pb1—O4	2.415 (11)	2.447 (15)
Pb1—O1	2.617 (14)	2.586 (14)
Pb1—X ⁱ	3.274 (3)	3.176 (12)
Pb1—X ⁱ	3.3287 (13)	3.237 (12)
Pb1—X ⁱⁱ	3.364 (3)	3.247 (12)
Pb2—O1	2.360 (13)	2.374 (14)
Pb2—O3 ⁱⁱⁱ	2.451 (18)	2.48 (2)
Pb2—O2	2.704 (19)	2.677 (13)
Pb2—X ²	3.3955 (17)	3.270 (12)
Pb2—X ^{iv}	3.415 (3)	3.276 (13)
Pb3—O4	2.555 (12)	2.544 (16)
Pb3—O1	2.617 (13)	2.600 (14)
Pb3—O4 ^v	2.788 (11)	2.777 (17)
Pb3—O3 ⁱⁱⁱ	2.995 (8)	2.986 (16)
Pb3—X ²	3.338 (21)	3.244 (12)
Te1—O1	1.989 (13)	1.938 (13)
Te1—O2	2.085 (17)	2.044 (12)
Te2—O3	1.874 (19)	1.84 (2)
Te2—O4	1.879 (11)	1.861 (16)

Symmetry codes: (i) $x + \frac{1}{2}, y - \frac{1}{2}, z$; (ii) $-x + \frac{1}{2}, -y + \frac{1}{2}, -z$; (iii) $-x + \frac{1}{2}, -y + \frac{1}{2}, -z + 1$; (iv) $x - \frac{1}{2}, y - \frac{1}{2}, z$; (v) $-x + \frac{1}{2}, y + \frac{1}{2}, -z + 1$.

Data collection: *APEX2* (Bruker, 2008); cell refinement: *SAINT* (Bruker, 2008); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ATOMS* (Dowty, 2006); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BR2129).

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supplementary materials

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Pb₃Te₂O₆Br₂

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Comment

Single crystals growth of Pb₃Te₂O₆Cl₂ and Pb₃Te₂O₆Br₂ was studied during a recent project intended to elaborate hydrothermal formation conditions of oxotelluraTe^{IV} compounds (Weil & Stöger, 2007; 2008*a,b*). Both Pb₃Te₂O₆X₂ (X = Br, Cl) compounds have been prepared previously by solid state techniques (Porter & Halasyamani, 2003). Whereas for the chloride compound a full structure analysis was undertaken at that time, for the isotopic bromide compound only lattice parameters were reported. Here we present details of the Pb₃Te₂O₆Br₂ structure determined from single-crystal X-ray data.

The asymmetric unit of the Pb₃Te₂O₆X₂ structure contains three Pb, two Te, two X and four O atoms. Except two of the O atoms, all other atoms are located on mirror planes. The Pb₃Te₂O₆X₂ structure type can be described as being built up from ²_∞[Pb₃Te₂O₆]²⁺ layers extending parallel to (20 $\bar{1}$) and X⁻ anions between the layers. The tellurium atoms in the cationic layer are surrounded by four (Te1) and three (Te2) oxygen atoms in distorted tetrahedral (Te1) and trigonal-pyramidal (Te2) environments, respectively. Under an additional contribution of the lone electron pairs to the stereochemistry of the two Te^{IV} atoms, the corresponding Ψ -oxopolyhedra can be considered as distorted TeO₄E square pyramids (Te1) and TeO₃E tetrahedra (E designates the lone electron pair). These kinds of TeO_x polyhedra are frequently observed for various oxotelluraTe^{IV} structures (Dolgikh, 1991; Zemmann, 1971). The Te1O₄ group forms Te₁₂O₆ dimers *via* edge-sharing, whereas the Te₂O₃ group is isolated in the layers. Both oxotelluraTe^{IV} units are surrounded by lead atoms. Four O atoms are bonded to Pb1, five O atoms to Pb2 and eight O atoms to Pb3. For both Pb₃Te₂O₆X₂ structures the respective Te—O and Pb—O distances are very similar (Table 1). The main difference between the Pb₃Te₂O₆X₂ structures pertains to the distances of the lead atoms to the X atoms that are situated between the cationic layers. As expected, the Pb—X distances are about 0.1 to 0.15 Å longer for the Br compound (Table 1). The layered character of the Pb₃Te₂O₆X₂ structure type with alternating layers parallel to (20 $\bar{1}$) is also reflected by the differences of the lattice parameters for the Br and the Cl analogues. While the lengths of the b-axes are very similar (5.6813 (3) (Br) *versus* 5.6295 (4) Å (Cl)), a and c differ notably (16.9151 (9) *versus* 16.4417 (11) Å, and 11.0623 (6) *versus* 10.8894 (7) Å) due to the different ionic radii of Br⁻ and Cl⁻. The coordination of the halogen atoms in the neighbouring anionic layer to the lead atoms augments the coordination polyhedra of Pb1 and Pb2 to an overall coordination of [Pb1O₄X₄], [Pb2O₅X₃] and [Pb3O₈X].

All TeO_x and PbO_xX_y polyhedra in the structure are highly irregular. The lone-pair electrons of both Te^{IV} and Pb^{II} atoms are stereochemically active and point towards the anionic halide layers (Fig. 1). Similar TeO_x and PbO_xX_y polyhedra are observed for the Pb₃TeO₄X₂ structures (Charkin *et al.*, 2006) that show a lower TeO₂ content in comparison with the Pb₃Te₂O₆X₂ structures.

Experimental

All chemicals used were of analytical grade and employed without further purification. 1 mmol PbBr_2 (0.367 g) and 1 mmol TeO_2 (0.160 g) were placed in a Teflon inlay that was filled with a hydrous NH_4OH solution (10%_w) to two-thirds of its volume. The inlay was placed in a steel autoclave and heated at 493 K for four weeks. The reaction product consisted of small colourless crystals of the title compound with rod-like habit and a maximum edge lengths of 0.2 mm. Experiments under similar conditions but with PbCl_2 instead of PbBr_2 led to crystals of the isotopic compound $\text{Pb}_3\text{Te}_2\text{O}_6\text{Cl}_2$.

Refinement

The structure was solved using direct methods. For better comparison with the isotopic structure of $\text{Pb}_3\text{Te}_2\text{O}_6\text{Cl}_2$ (Porter & Halasyamani, 2003), the atomic coordinates of the Cl analogue were taken as starting parameters for refinement. The highest remaining peak in the final difference Fourier map is 0.80 Å from Pb2 and the deepest hole is 0.57 Å from the same atom.

Lattice parameters of $\text{Pb}_3\text{Te}_2\text{O}_6\text{Br}_2$ based on the present single-crystal study agree reasonably well with those of the powder diffraction data provided by Porter & Halasyamani (2003): $a = 16.8911$ (8), $b = 5.6804$ (2), $c = 11.0418$ (5) Å, $\beta = 104.253$ (2)°.

Figures

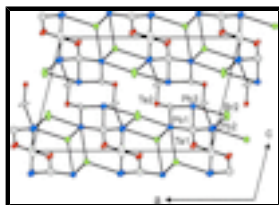


Fig. 1. The crystal structure of $\text{Pb}_3\text{Te}_2\text{O}_6\text{Br}_2$ in projection along [010]. Te atoms are given as red, Pb as blue, O atoms as white and Br atoms as green ellipsoids at the 74% probability level.

trilead(II) bis[telluraTe^{IV}] dibromide

Crystal data

$\text{Pb}_3\text{Te}_2\text{O}_6\text{Br}_2$

$M_r = 1132.59$

Monoclinic, $C2/m$

Hall symbol: $-C 2y$

$a = 16.9151$ (9) Å

$b = 5.6813$ (3) Å

$c = 11.0623$ (6) Å

$\beta = 104.046$ (1)°

$V = 1031.3$ (1) Å³

$Z = 4$

$F(000) = 1872$

$D_x = 7.295$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 4771 reflections

$\theta = 2.5\text{--}32.2^\circ$

$\mu = 62.14$ mm⁻¹

$T = 296$ K

Rod, colourless

$0.18 \times 0.10 \times 0.04$ mm

Data collection

Bruker APEXII CCD

1371 independent reflections

diffractometer
 Radiation source: fine-focus sealed tube 1338 reflections with $I > 2\sigma(I)$
 graphite $R_{\text{int}} = 0.030$
 ω - and φ -scans $\theta_{\text{max}} = 27.9^\circ$, $\theta_{\text{min}} = 1.9^\circ$
 Absorption correction: numerical
 (*HABITUS*; Herrendorf, 1997) $h = -22 \rightarrow 15$
 $T_{\text{min}} = 0.06$, $T_{\text{max}} = 0.41$ $k = -7 \rightarrow 7$
 3758 measured reflections $l = -13 \rightarrow 14$

Refinement

Refinement on F^2 Primary atom site location: structure-invariant direct methods
 Least-squares matrix: full Secondary atom site location: difference Fourier map
 $R[F^2 > 2\sigma(F^2)] = 0.048$ $w = 1/[\sigma^2(F_o^2) + (0.0166P)^2 + 250.8836P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $wR(F^2) = 0.107$ $(\Delta/\sigma)_{\text{max}} < 0.001$
 $S = 1.15$ $\Delta\rho_{\text{max}} = 7.41 \text{ e } \text{\AA}^{-3}$
 1371 reflections $\Delta\rho_{\text{min}} = -6.41 \text{ e } \text{\AA}^{-3}$
 73 parameters Extinction correction: *SHELXL97* (Sheldrick, 2008),
 $F_c^* = kFc[1 + 0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
 0 restraints Extinction coefficient: 0.00021 (2)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Pb1	0.26158 (6)	0.0000	0.21028 (8)	0.0178 (2)
Pb2	0.02616 (6)	0.0000	0.19788 (9)	0.0250 (3)
Pb3	0.16368 (6)	0.5000	0.39320 (8)	0.0217 (3)
Te1	0.10510 (10)	0.5000	0.04764 (16)	0.0233 (4)
Te2	0.37019 (9)	0.5000	0.41467 (14)	0.0141 (3)
Br1	0.31858 (15)	0.5000	0.0993 (2)	0.0205 (5)
Br2	-0.03969 (16)	0.5000	0.3103 (3)	0.0313 (6)
O1	0.1315 (8)	0.263 (3)	0.1840 (11)	0.028 (3)
O2	0.0000	0.294 (5)	0.0000	0.059 (7)
O3	0.3884 (15)	0.5000	0.5887 (17)	0.040 (5)

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O4 0.2909 (7) 0.262 (2) 0.3877 (10) 0.017 (2)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Pb1	0.0224 (5)	0.0150 (4)	0.0161 (4)	0.000	0.0048 (3)	0.000
Pb2	0.0258 (5)	0.0204 (5)	0.0253 (5)	0.000	-0.0004 (4)	0.000
Pb3	0.0190 (4)	0.0261 (5)	0.0191 (4)	0.000	0.0026 (3)	0.000
Te1	0.0224 (8)	0.0186 (8)	0.0236 (8)	0.000	-0.0046 (6)	0.000
Te2	0.0117 (6)	0.0124 (7)	0.0177 (7)	0.000	0.0025 (5)	0.000
Br1	0.0208 (11)	0.0220 (11)	0.0183 (11)	0.000	0.0039 (9)	0.000
Br2	0.0202 (12)	0.0236 (13)	0.0525 (18)	0.000	0.0131 (12)	0.000
O1	0.032 (7)	0.026 (7)	0.023 (6)	-0.010 (6)	0.005 (5)	0.002 (5)
O2	0.031 (11)	0.062 (18)	0.09 (2)	0.000	0.028 (13)	0.000
O3	0.051 (13)	0.046 (14)	0.011 (8)	0.000	-0.014 (8)	0.000
O4	0.019 (5)	0.016 (6)	0.018 (5)	-0.007 (5)	0.009 (4)	-0.001 (4)

Geometric parameters (\AA , $^\circ$)

Pb1—O4	2.415 (11)	Pb3—O4 ^{viii}	2.555 (12)
Pb1—O4 ⁱ	2.415 (11)	Pb3—O4	2.555 (12)
Pb1—O1 ⁱ	2.617 (14)	Pb3—O1	2.617 (13)
Pb1—O1	2.617 (14)	Pb3—O1 ^{viii}	2.617 (13)
Pb1—Br2 ⁱⁱ	3.274 (3)	Pb3—O4 ^{ix}	2.788 (11)
Pb1—Br1 ⁱⁱⁱ	3.3287 (13)	Pb3—O4 ^v	2.788 (11)
Pb1—Br1	3.3287 (13)	Pb3—O3 ^v	2.995 (8)
Pb1—Br1 ^{iv}	3.364 (3)	Pb3—O3 ^x	2.995 (8)
Pb2—O1	2.360 (13)	Pb3—Te2	3.4432 (17)
Pb2—O1 ⁱ	2.360 (13)	Te1—O1 ^{viii}	1.989 (13)
Pb2—O3 ^v	2.451 (18)	Te1—O1	1.989 (13)
Pb2—O2 ^{vi}	2.704 (19)	Te1—O2	2.085 (17)
Pb2—O2	2.704 (19)	Te1—O2 ^{xi}	2.085 (17)
Pb2—Br2	3.3955 (17)	Te2—O3	1.874 (19)
Pb2—Br2 ⁱⁱⁱ	3.3955 (17)	Te2—O4	1.879 (11)
Pb2—Br1 ^{vii}	3.415 (3)	Te2—O4 ^{viii}	1.879 (11)
O4—Pb1—O4 ⁱ	76.0 (5)	O1—Te1—O2	80.4 (6)
O4—Pb1—O1 ⁱ	116.3 (4)	O1 ^{viii} —Te1—O2 ^{xi}	80.4 (6)
O4 ⁱ —Pb1—O1 ⁱ	74.9 (4)	O1—Te1—O2 ^{xi}	126.3 (5)
O4—Pb1—O1	74.9 (4)	O2—Te1—O2 ^{xi}	68.2 (14)
O4 ⁱ —Pb1—O1	116.3 (4)	O3—Te2—O4	95.5 (6)
O1 ⁱ —Pb1—O1	69.7 (6)	O3—Te2—O4 ^{viii}	95.5 (6)
O1—Pb2—O1 ⁱ	78.6 (7)	O4—Te2—O4 ^{viii}	92.3 (7)
O1—Pb2—O3 ^v	77.6 (5)	Te1—O1—Pb2	116.3 (6)
O1 ⁱ —Pb2—O3 ^v	77.6 (5)	Te1—O1—Pb1	119.9 (6)

O1—Pb2—O2 ^{vi}	108.5 (4)	Pb2—O1—Pb1	105.0 (5)
O1 ⁱ —Pb2—O2 ^{vi}	62.2 (4)	Te1—O1—Pb3	106.4 (6)
O3 ^v —Pb2—O2 ^{vi}	136.2 (5)	Pb2—O1—Pb3	105.4 (5)
O1—Pb2—O2	62.2 (4)	Pb1—O1—Pb3	101.9 (4)
O1 ⁱ —Pb2—O2	108.5 (4)	Te1—O2—Te1 ^{xi}	111.8 (14)
O3 ^v —Pb2—O2	136.2 (5)	Te1—O2—Pb2 ^{vi}	120.87 (15)
O2 ^{vi} —Pb2—O2	76.4 (10)	Te1 ^{xi} —O2—Pb2 ^{vi}	100.36 (16)
O4 ^{viii} —Pb3—O4	64.0 (5)	Te1—O2—Pb2	100.36 (16)
O4 ^{viii} —Pb3—O1	104.3 (4)	Te1 ^{xi} —O2—Pb2	120.87 (15)
O4—Pb3—O1	72.7 (4)	Pb2 ^{vi} —O2—Pb2	103.6 (10)
O4 ^{viii} —Pb3—O1 ^{viii}	72.7 (4)	Te2—O3—Pb2 ^v	154.3 (13)
O4—Pb3—O1 ^{viii}	104.3 (4)	Te2—O4—Pb1	124.8 (5)
O1—Pb3—O1 ^{viii}	61.9 (6)	Te2—O4—Pb3	100.8 (5)
O1 ^{viii} —Te1—O1	85.1 (8)	Pb1—O4—Pb3	109.7 (4)
O1 ^{viii} —Te1—O2	126.3 (5)		

Symmetry codes: (i) $x, -y, z$; (ii) $x+1/2, y-1/2, z$; (iii) $x, y-1, z$; (iv) $-x+1/2, -y+1/2, -z$; (v) $-x+1/2, -y+1/2, -z+1$; (vi) $-x, -y, -z$; (vii) $x-1/2, y-1/2, z$; (viii) $x, -y+1, z$; (ix) $-x+1/2, y+1/2, -z+1$; (x) $-x+1/2, -y+3/2, -z+1$; (xi) $-x, -y+1, -z$.

Table 1

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Pb3—O4 ^v	2.788 (11)	2.777 (17)
Pb3—O3 ⁱⁱⁱ	2.995 (8)	2.986 (16)
Pb3—X2	3.338 (21)	3.244 (12)
Te1—O1	1.989 (13)	1.938 (13)
Te1—O2	2.085 (17)	2.044 (12)
Te2—O3	1.874 (19)	1.84 (2)
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Symmetry codes: (i) $x+1/2, y-1/2, z$; (ii) $-x+1/2, -y+1/2, -z$; (iii) $-x+1/2, -y+1/2, -z+1$; (iv) $x-1/2, y-1/2, z$; (v) $-x+1/2, y+1/2, -z+1$.

Fig. 1

